In the space provided, write the letter of the description that best matches the term or phrase.

1. suspension	**a.** in a solution, the substance in which the solute is dissolved
2. colloid	**b.** describes a solution in which the solvent is water
3. solution	**c.** a mixture in which large particles are temporarily suspended throughout a liquid or gas
4. solvent	**d.** in a solution, the substance dissolved in the solvent
5. solute	**e.** a stable, homogeneous mixture
6. aqueous	**f.** a mixture of very small particles stably suspended in liquid, solid, or gas

Mark each mixture below *H* if it is homogeneous and *Ht* if it is heterogeneous

_____ **7.** milk
_____ **8.** gasoline
_____ **9.** muddy water
_____**10.** steel
_____**11.** coffee
_____**12.** ink
_____**13.** crude oil
_____**14.** bronze
_____**15.** brass
_____**16.** sea water

Complete each statement with the correct term or phrase.

17. All parts of a_____ mixture have the same composition.

18. **_____**mixtures are not uniform in composition.

19. Any mixture that is heterogeneous on a microscopic level is a_____
.

20. A _____ is a mixture in which particles of the mixture are evenly dispersed throughout a liquid or gas.

23. **_____**are intermediate between suspensions and solutions.

Naming and writing names for Ionic Compounds

24. All of the following are homogeneous mixtures except
a. tomato soup.
b. a sugar-water solution.
c. gasoline.
d. a salt-water solution.

25. Which of the following is a colloid?
a. water
b. milk
c. soil
d. concrete

26. A mixture that appears to be uniform while being stirred but separates into different phases when agitation ceases is a
a. solvent.
b. colloid.
c. suspension.
d. solute.

27. Which of the following is not a solute-solvent combination?
a. gas-gas
b. gas-liquid
c. gas-solid
d. liquid-solid

28. Carbon dioxide in air is an example of which solute-solvent combination?
a. gas-liquid
b. liquid-gas
c. liquid-liquid
d. gas-gas

29. Sugar in water is an example of which solute-solvent combination?
a. gas-liquid
b. liquid-liquid
c. solid-liquid
d. liquid-solid

30. Comparing the size of the particles in a solution and in a colloid, the particle size in the solution is
a. smaller.
b. larger.
c. the same size as the particle size in the colloid.
d. dependent on the colloid and the solution.

31. You have a mixture consisting of salt and iron filings. How might you separate them?
a. Use a magnet to attract the iron filings.

b. Dissolve the salt in water, then filter and evaporate.
c. Decant the salt. d. Both (a) and (b)

Naming Ionic Compounds

Write the names for the following ionic compounds	Write the formulas for the following ionic compounds
1. NaCl 2. $MgCl_2$ 3. Al_2O_3 4. $CaSO_4$ 5. NH_4SO_3 6. K_2CO3	8. Potassium chloride 9. Sodium sulfide 10. Aluminum phosphate 11. Ammonium sulphate 12. Sodium oxide 13. Magnesium oxide

Naming Molecular Compounds

Prefixes

1- Mono	2-Di	3-Tri	4-Tetra	5-Penta

Write the names for the following molecular compounds	Write the formulas for the following ionic compounds
1. HCl 2. ICl_2 3. P_2O_5 4. PCl_3 5. N_2F_2 6. CCl_4 7. I_2 8. Cl_2 9. O_2	7. Carbon dioxide 8. Dinitrogen Pentoxide 9. Phosphorous Pentachloride 10. Bromine 11. Sodium oxide 12. Magnesium oxide 13. Sulfur dioxide 14. Carbon tetrabromide